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Kinetics of Electrochemical Reduction of Propyl Halides in Aprotic Media

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Abstract

The free energies of activation of reduction of n-propyl halides (Hlg = Cl, Br, I) on mercury in aprotic media (acetone, acetonitrile, dimethylformamide) vary within the range 0.199-0.260 eV. The reduction proceeds by an activation mechanism with simultaneous C-Hlg bond cleavage. A uniform transiton state is realized in all the solvents studied.